

Assam Board Class 10 General Science Question Paper 2026

Time Allowed :3 Hours

Maximum Marks :100

Total questions :24

General Instructions

Read the following instructions very carefully and strictly follow them:

1. The paper is divided into two sections – Section A (Compulsory) and Section B (Elective).
2. Section A is compulsory for all candidates and generally includes objective-type questions, short answer questions, and long answer questions from the prescribed syllabus.
3. In Section A, candidates are required to answer all questions. The questions will cover topics from ancient, medieval, and modern history as prescribed by the syllabus.
4. Section B consists of elective questions. Candidates are required to attempt questions from the chosen topic according to the provided options.
5. The questions in Section A will be in the form of multiple-choice, short answer, and essay-type questions.
6. Answers to all questions must be written in neat and legible handwriting. Candidates must adhere strictly to the word limit mentioned in the questions.
7. Use of unfair means or electronic devices during the examination is strictly prohibited.
8. Candidates must ensure that they write their answers in the correct format, following the instructions given for each section.

1. Identify the product which represents the solid state in the above reaction.

- (A) Barium chloride
 - (B) Barium sulphate
 - (C) Sodium chloride
 - (D) Sodium sulphate
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2. The colour of the solution observed after 30 minutes of placing zinc metal to copper sulphate solution is

- (A) Blue
 - (B) Colourless
 - (C) Dirty green
 - (D) Reddish Brown
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3. In the redox reaction



- (A) MnO_2 is reduced to MnCl_2 & HCl is oxidized to H_2O
 - (B) MnO_2 is reduced to MnCl_2 & HCl is oxidised to Cl_2
 - (C) MnO_2 is oxidized to MnCl_2 & HCl is reduced to Cl_2
 - (D) MnO_2 is oxidized to MnCl_2 & HCl is Reduced to H_2O
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4. On placing a copper coin in a test tube containing green ferrous sulphate solution, it will be observed that the ferrous sulphate solution

- (A) turns blue, and a grey substance is deposited on the copper coin.
 - (B) turns colourless and a grey substance is deposited on the copper coin.
 - (C) turns colourless and a reddish-brown substance is deposited on the copper coin.
 - (D) remains green with no change in the copper coin.
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5. Reema took 5ml of Lead Nitrate solution in a beaker and added approximately 4ml of Potassium Iodide solution to it. What would she observe?

- (A) The solution turned red.
 - (B) Yellow precipitate was formed.
 - (C) White precipitate was formed.
 - (D) The reaction mixture became hot.
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6. Why is it important to balance a skeletal chemical equation?

- (A) To verify law of conservation of energy.
 - (B) To verify the law of constant proportion.
 - (C) To verify the law of conservation of mass.
 - (D) To verify the law of conservation of momentum.
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7. Which of the following correctly represents a balanced chemical equation?

- (A) $\text{Fe(s)} + 4\text{H}_2\text{O(g)} \rightarrow \text{Fe}_3\text{O}_4\text{(s)} + 4\text{H}_2\text{(g)}$
 - (B) $3\text{Fe(s)} + 4\text{H}_2\text{O(g)} \rightarrow \text{Fe}_3\text{O}_4\text{(s)} + 4\text{H}_2\text{(g)}$
 - (C) $3\text{Fe(s)} + \text{H}_2\text{O(g)} \rightarrow \text{Fe}_3\text{O}_4\text{(s)} + \text{H}_2\text{(g)}$
 - (D) $3\text{Fe(s)} + 4\text{H}_2\text{O(g)} \rightarrow \text{Fe}_3\text{O}_4\text{(s)} + \text{H}_2\text{(g)}$
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8. The chemical reaction between copper and oxygen can be categorized as:

- (A) Displacement reaction
 - (B) Decomposition reaction
 - (C) Combination reaction
 - (D) Double displacement reaction
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9. Anita added a drop each of diluted acetic acid and diluted hydrochloric acid on pH paper and compared the colors. Which of the following is the correct conclusion?

- (A) pH of acetic acid is more than that of hydrochloric acid.
- (B) pH of acetic acid is less than that of hydrochloric acid.
- (C) Acetic acid dissociates completely in aqueous solution.

(D) Acetic acid is a strong acid.

10. An element 'M' has 50% of the electrons filled in the 3rd shell as in the 2nd shell.

The atomic number of 'M' is:

- (A) 10
 - (B) 12
 - (C) 14
 - (D) 18
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11. Generally food is broken and absorbed within the body of organisms. In which of the following organisms is it done outside the body?

- (A) Amoeba
 - (B) Mushroom
 - (C) Paramecium
 - (D) Lice
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12. Opening and closing of stomatal pore depends on:

- (A) Atmospheric temperature
 - (B) oxygen concentration around stomata
 - (C) carbon dioxide concentration around stomata
 - (D) water content in the guard cells
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13. Given below are the functions of some parts of human circulatory system. Identify the correct match.

- (A) Pulmonary vein - takes oxygenated blood from body parts to heart
- (B) Artery - takes oxygenated blood from heart to lung
- (C) Dorsal aorta - takes deoxygenated blood from heart to body parts
- (D) Vena cava - takes deoxygenated blood from body parts to right atrium

14. What happens when right and left ventricle contract during pumping of blood by human heart?

- (A) Blood transferred to the right ventricle and left ventricle simultaneously.
- (B) Blood is transferred to lungs for oxygenation and is pumped into various organs simultaneously.
- (C) Blood transferred to the right atrium and left atrium simultaneously.
- (D) Blood is received from lungs after oxygenation and is received from various organs of the body.

15. This experimental set up is used to prove essentiality of which of the following requirements of photosynthesis?

- (A) Chlorophyll
- (B) Oxygen
- (C) Carbon dioxide
- (D) Sunlight

16. The function of KOH is to absorb

- (A) Oxygen.
- (B) Carbon dioxide.
- (C) Moisture.
- (D) Sunlight.

17. Rays from Sun converge at a point 15 cm in front of a concave mirror. Where should an object be placed so that size of its image is equal to the size of the object?

- (A) 30 cm in front of the mirror
- (B) 15 cm in front of the mirror
- (C) Between 15 cm and 30 cm in front of the mirror

(D) More than 30 cm in front of the mirror

18. Nalin draws a ray diagram for an object in front of a concave mirror. She draws a ray starting from the top of the object and falling on the mirror perpendicularly. The ray after reflection will

- (A) pass through focus.
 - (B) pass through pole.
 - (C) pass through the centre of curvature.
 - (D) pass through any point on the principal axis.
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19. If the refractive index of water with respect to air is 1.33 and of that of glass with respect to air is 1.5 then

- (A) water is optically denser than glass.
 - (B) air is optically densest of all the three media.
 - (C) air's optical density is between glass and air.
 - (D) glass is optically denser than water.
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20. Chlorine gas was prepared using electrolysis of brine solution. Write the chemical equation to represent the change. Identify the other products formed in the process and give one application of each.

21. Chlorine gas was prepared using electrolysis of brine solution. Write the chemical equation to represent the change. Identify the other products formed in the process and give one application of each.

22. Rehmat classified the reaction between Methane and Chlorine in presence of sunlight as a substitution reaction. Support Rehmat's view with suitable justification and illustrate the reaction with the help of a balanced chemical equation.

23. A lot of waste is generated in neighborhood. However, almost all of it is biodegradable. What impact will it have on the environment or human health?

24. Why is damage to the ozone layer a cause for concern? What are its causes and what steps are being taken to limit this damage?
