

NEET Re-Exam 2026 Chemistry

Question Paper

Conducted by National Testing Agency (NTA)

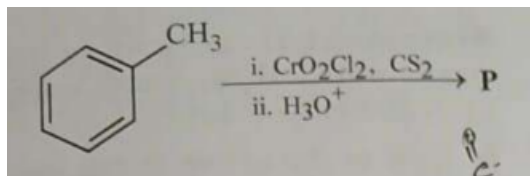


General Instructions

- (i) The test is of 3 hours and 15 minutes duration.
- (ii) This test paper consists of 180 questions. The maximum marks are 720.
- (iii) Physics and Chemistry contains 45 questions each and Biology (Botany and Zoology) contains 90 questions.
- (iv) Each question carries +4 marks for correct answer and -1 mark for wrong answer.

Chemistry

46. Consider the following reaction, and choose the correct option.



- (A) Compound P is obtained by the hydrogenation of benzoyl chloride with Pd on BaSO₄.
- (B) On treating compound P with saturated NaHCO₃ solution, brisk effervescence is observed.
- (C) Compound P can be prepared by treating benzene with anhydrous AlCl₃ and CH₃COCl.
- (D) On treatment with bromine water, compound P gives a white precipitate.

47. The formula of tetraammineaquachloridocobalt(III) chloride is

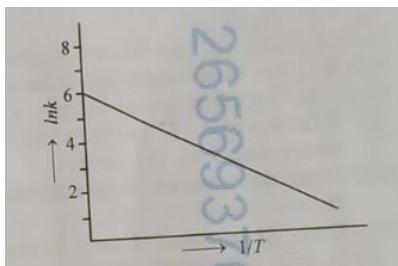
- (A) [Co(NH₃)₄(H₂O)Cl]Cl₂
- (B) [Co(NH₃)₄Cl₂] × H₂O
- (C) [Co(NH₃)₄]Cl₃ × H₂O
- (D) [Co(NH₃)₄(H₂O)Cl]Cl

48. The lanthanide ion having four unpaired electrons is

(Given : Atomic numbers of Ce = 58, Nd = 60, Tb = 65 and Ho = 67)

- (1) Ho^{3+}
 - (2) Nd^{3+}
 - (3) Ce^{3+}
 - (4) Tb^{3+}
-

49. For an elementary chemical reaction, the Arrhenius plot is given below.



If the energy of activation is 6.64 kJ mol^{-1} and $R = 8.3 \text{ J K}^{-1} \text{ mol}^{-1}$, the temperature at which the rate constant becomes $e^2 \text{ min}^{-1}$, is

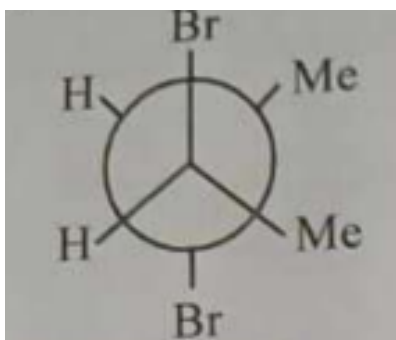
- (1) 250 K
 - (2) 125 K
 - (3) 150 K
 - (4) 200 K
-

50. The green paramagnetic species formed by heating KMnO_4 at 513 K is

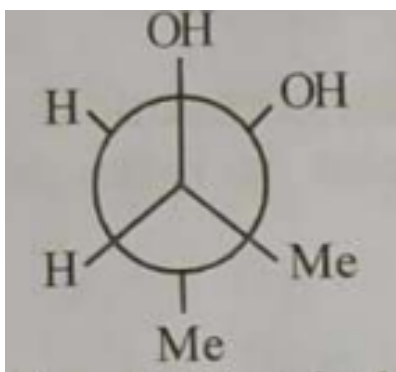
- (1) KO_2
 - (2) K_2MnO_4
 - (3) Mn_3O_4
 - (4) MnO
-

51. Given below are two statements:

Statement I: *trans*-But-2-ene upon treatment with Br_2 in CCl_4 gives the following product:



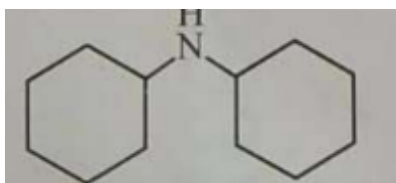
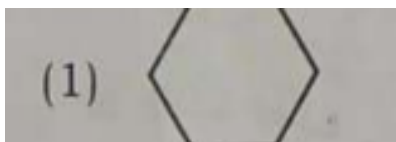
Statement II: *cis*-But-2-ene upon treatment with alkaline KMnO_4 gives the following product:

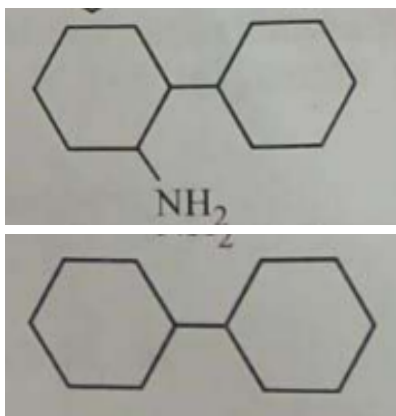


In the light of the above statements, choose the *most appropriate* answer from the options given below.

- (1) Statement I is incorrect but Statement II is correct
- (2) Both Statement I and Statement II are correct
- (3) Both Statement I and Statement II are incorrect
- (4) Statement I is correct but Statement II is incorrect

52. One of the products formed in the following reaction is





- (1) FigA
- (2) FigB
- (3) FigC
- (4) FigD

53. Given below are two statements:

Statement-I : Heating NaCl with concentrated H_2SO_4 and MnO_2 results in oxidation of Mn.

Statement-II : Heating NaI with concentrated H_2SO_4 and MnO_2 results in reduction of Mn.

In light of the above statements, choose the *most appropriate* answer from the options given below:

- (1) Statement-I is incorrect but Statement-II is correct
- (2) Both Statement-I and Statement-II are correct
- (3) Both Statement-I and Statement-II are incorrect
- (4) Statement-I is correct but Statement-II is incorrect

54. Among the following options, the correct trend in the electron gain enthalpy is

- (1) $\text{I} > \text{Br} > \text{Cl} > \text{F}$
- (2) $\text{F} > \text{Cl} > \text{Br} > \text{I}$
- (3) $\text{Br} > \text{Cl} > \text{F} > \text{I}$
- (4) $\text{Cl} > \text{F} > \text{Br} > \text{I}$

55. Given below are two statements:

Statement-I : $[\text{Fe}(\text{ox})_3]^{3-}$ is chiral.

Statement-II : $\text{trans-}[\text{Cr}(\text{H}_2\text{O})_2(\text{ox})_2]^-$ is chiral.

(Given : $\text{oxH}_2 = \text{HOOC} - \text{COOH}$)

In light of the above statements, choose the *most appropriate* answer from the options given below:

- (1) Statement-I is incorrect but Statement-II is correct
 - (2) Both Statement-I and Statement-II are correct
 - (3) Both Statement-I and Statement-II are incorrect
 - (4) Statement-I is correct but Statement-II is incorrect
-

56. The correct statement about peptides and proteins is

- (1) In α -helices, the polypeptide chain is twisted into a left-handed screw (helix) through intramolecular hydrogen bonds.
 - (2) Tertiary structure of proteins has two or more polypeptide subunits.
 - (3) Only the proteins having a quaternary structure are biologically active.
 - (4) In β -pleated sheet structures, peptide chains are held together by intermolecular hydrogen bonds.
-

57. Given below are two statements:

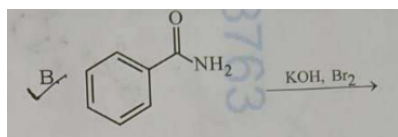
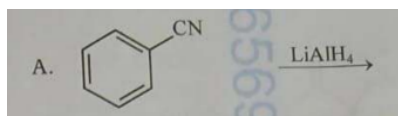
Statement-I : Oxidation of *p*-nitrotoluene with acidic KMnO_4 gives an acid that is stronger than benzoic acid.

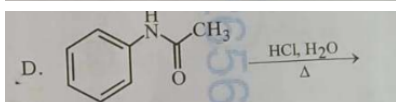
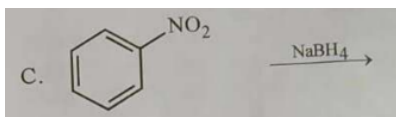
Statement-II : Reduction of *p*-nitrotoluene with Sn/HCl followed by neutralization gives an amine that is more basic than aniline.

In light of the above statements, choose the *most appropriate* answer from the options given below.

- (1) Statement-I is incorrect but Statement-II is correct.
 - (2) Both Statement-I and Statement-II are correct.
 - (3) Both Statement-I and Statement-II are incorrect.
 - (4) Statement-I is correct but Statement-II is incorrect.
-

58. Identify the reactions which give aniline as the major product.



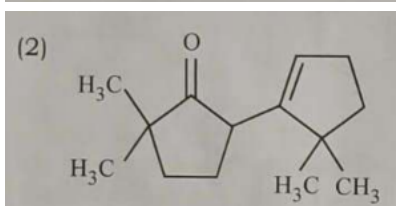
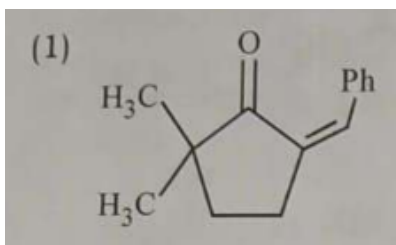
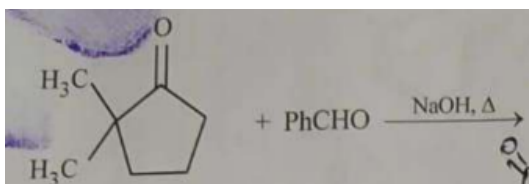


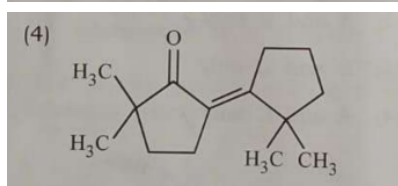
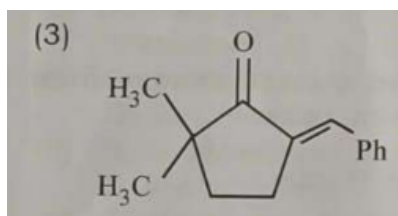
- (1) FigA
- (2) FigB
- (3) FigC
- (4) FigD

59. Two moles of an ideal gas undergo free expansion from 10 L to 100 L at 300 K. The values of ΔS_{system} and $\Delta S_{\text{surroundings}}$ are (R is universal gas constant)

- (1) $\Delta S_{\text{system}} = 4.606 R$; $\Delta S_{\text{surroundings}} = 0$
- (2) $\Delta S_{\text{system}} = 0$; $\Delta S_{\text{surroundings}} = 0$
- (3) $\Delta S_{\text{system}} = 4.606 R$; $\Delta S_{\text{surroundings}} = -4.606 R$
- (4) $\Delta S_{\text{system}} = 0$; $\Delta S_{\text{surroundings}} = 4.606 R$

60. The compound that CANNOT be obtained from the aldol condensation reaction shown below, is





- (1) FigA
(2) FigB
(3) FigC
(4) FigD

61. The complex which has *facial* and *meridional* isomers is

(Given : py = pyridine and en = $\text{H}_2\text{N} - \text{CH}_2 - \text{CH}_2 - \text{NH}_2$)

- (1) $[\text{Ni}(\text{en})_2(\text{H}_2\text{O})_2]^{2+}$
(2) $[\text{Cr}(\text{py})_3(\text{Cl})_3]$
(3) $[\text{Cr}(\text{H}_2\text{O})_6]^{3+}$
(4) $[\text{Co}(\text{NH}_3)_4(\text{H}_2\text{O})_2]^{3+}$

62. The numbers 17.0145 and 21.0235 were rounded to three figures after the decimal point. The resulting numbers, respectively, are

- (1) 17.015 and 21.024
(2) 17.014 and 21.023
(3) 17.015 and 21.023
(4) 17.014 and 21.024

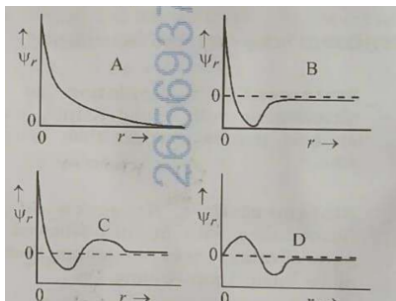
63. The amount of carbon dioxide evolved upon complete combustion of 116 g of *n*-butane is
(Given: atomic mass in amu H = 1, C = 12 and O = 16)

- (1) 362 g
(2) 352 g

(3) 322 g

(4) 176 g

64. Consider the following schematic plots of orbital wavefunction (ψ_r) against distance (r) from the nucleus.



The figure representing two radial nodes in the orbital is

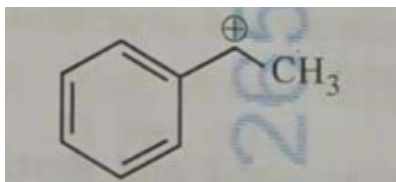
(1) D

(2) A

(3) B

(4) C

65. The following carbocation is stabilized by the interaction of the empty p orbital with



(1) empty σ^* and empty π^* orbitals

(2) filled σ and filled π orbitals

(3) empty σ and empty π^* orbitals

(4) empty σ^* and filled π orbitals

66. A 1:3 electrolyte in an aqueous solution is

(1) $[\text{Co}(\text{NH}_3)_3(\text{NO}_2)_3]$

(2) $[\text{CoCl}_2(\text{NH}_3)_4]\text{Cl}$

(3) $[\text{CoCl}(\text{NH}_3)_5]\text{Cl}_2$

(4) $[\text{Co}(\text{NH}_3)_6]\text{Cl}_3$

67. The standard electrode potential (E°) for the half-cell reaction $\text{Fe}^{3+} + e^- \rightarrow \text{Fe}^{2+}$ at 298 K is (Given: $E^\circ(\text{Fe}^{3+}/\text{Fe}) = -0.04 \text{ V}$ and $E^\circ(\text{Fe}^{2+}/\text{Fe}) = -0.44 \text{ V}$ at 298 K)

- (1) +0.92 V
 - (2) +0.40 V
 - (3) +0.76 V
 - (4) -0.48 V
-

68. In potash alum, the ratio of K^+ and SO_4^{2-} ions is

- (1) 3 : 2
 - (2) 1 : 2
 - (3) 2 : 1
 - (4) 2 : 3
-

69. Consider the following statements about the solutions formed by mixing two liquids.

- A. An ideal solution thus formed obeys Raoult's law throughout the composition range.
- B. Mixture of chloroform and acetone shows negative deviation from Raoult's law.
- C. Mixture of aniline and phenol shows positive deviation from Raoult's law.

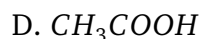
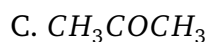
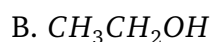
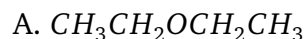
Select the correct option:

- (1) A and C only
 - (2) A and B only
 - (3) B and C only
 - (4) A only
-

70. For a salt XY, which is a strong electrolyte, the plot of Λ_m versus \sqrt{c} has a slope of $-90.0 \text{ S cm}^2 \text{ mol}^{-3/2} \text{ L}^{1/2}$ at 298 K. At 0.01 M concentration of XY, the value of Λ_m is $145.0 \text{ S cm}^2 \text{ mol}^{-1}$. The limiting molar conductivity of Y^- ion ($\lambda_{\text{Y}^-}^\circ$, in $\text{S cm}^2 \text{ mol}^{-1}$) at 298 K will be (Given: $\lambda_{\text{X}^+}^\circ = 74.0 \text{ S cm}^2 \text{ mol}^{-1}$)

- (1) 76.0
 - (2) 80.0
 - (3) 100.0
 - (4) 90.0
-

71. Arrange the following compounds in the increasing order of polarity:



Choose the correct answer from the options given below:

- (A) $A < C < B < D$
 - (B) $A < B < C < D$
 - (C) $C < A < D < B$
 - (D) $C < A < B < D$
-

72. According to crystal field theory, the correct order of ligands with respect to their decreasing order of field strength is

- (A) $Cl^- > NH_3 > H_2O > CO$
 - (B) $CO > NH_3 > H_2O > Cl^-$
 - (C) $CO > H_2O > NH_3 > Cl^-$
 - (D) $Cl^- > H_2O > NH_3 > CO$
-

73. The amino acid that gives a red-blood colour on treating its sodium fusion extract with sodium nitroprusside is

- (A) serine

- (B) leucine
 - (C) threonine
 - (D) methionine
-

74. In an acidic medium, 10 mL of 0.25 M oxalic acid is titrated with $KMnO_4$ solution. If the volume of $KMnO_4$ solution required to reach the end point is 10 mL, the strength of the $KMnO_4$ solution is

- (A) 0.15 M
 - (B) 0.10 M
 - (C) 0.20 M
 - (D) 0.25 M
-

75. The correct statement is

- (A) Aluminium has five valence orbitals.
 - (B) Boron has a maximum covalency of four.
 - (C) Beryllium has three valence orbitals.
 - (D) Magnesium has a maximum covalency of four.
-

76. Among the following, the compound having conjugated double bonds is

- (A) hepta-1,6-diene
 - (B) hepta-1,3-diene
 - (C) hepta-1,4-diene
 - (D) hepta-1,5-diene
-

77. For a zero-order reaction, where $k = 1.0 \text{ mol L}^{-1} \text{ min}^{-1}$. If the initial concentration of A is 2 M, then the time taken for completion of 75% of the reaction will be

- (A) 2.0 min
 - (B) 1.5 min
 - (C) 0.75 min
 - (D) 1.0 min
-

78. The correct order of solubility of the given salts in water at 298 K is

Salt	K_{sp} at 298 K
AgBr	5.0×10^{-13}
Zn(OH) ₂	1.0×10^{-15}
Hg ₂ Cl ₂	1.3×10^{-18}

- (A) $Zn(OH)_2 > AgBr > Hg_2Cl_2$
(B) $Hg_2Cl_2 > Zn(OH)_2 > AgBr$
(C) $AgBr > Zn(OH)_2 > Hg_2Cl_2$
(D) $Hg_2Cl_2 > AgBr > Zn(OH)_2$
-

79. The correct decreasing order of oxidation state of the underlined atom in each molecule is

- (A) $P_4O_{10} > Cl_2O_7 > AlH_3$
(B) $P_4O_{10} > SO_3 > H_2O$
(C) $N_2O_5 > Al_2O_3 > H_2S$
(D) $PbO_2 > N_2O_3 > SO_3$
-

80. Consider the reversible processes for 1.0 mol of an ideal gas as shown in the figure. Processes 2 and 4 are adiabatic. w_1, w_2, w_3 and w_4 represent work done (in calories) in processes 1, 2, 3 and 4, respectively. ΔU_2 and ΔU_4 are changes in internal energy for processes 2 and 4, respectively. [Use $R = 2 \text{ cal K}^{-1} \text{ mol}^{-1}$] The correct option is

- (A) $w_1 + w_2 + w_3 + w_4 = 0$
(B) $w_1 + w_3 = -2T_1 \ln\left(\frac{V_2}{V_1}\right) - 2T_2 \ln\left(\frac{V_4}{V_3}\right)$
(C) $w_2 + w_4 = \Delta U_2 - \Delta U_4$
(D) $w_1 + w_2 = 2T_1 \ln\left(\frac{V_2}{V_1}\right)$
-

81. Assertion A : For an ideal solution formed by mixing liquids P and Q, $\Delta_{mix}H = 0$ and $\Delta_{mix}V = 0$.

Reason R : No interactions occur between P and Q.

In the light of the above statements, choose the most appropriate answer from the options given below.

- (A) A is not correct but R is correct.
(B) Both A and R are correct and R is the correct explanation of A.
(C) Both A and R are correct but R is NOT the correct explanation of A.
(D) A is correct but R is not correct.
-

82. Among the species given below, the spin-only magnetic moment is highest for

(Given: Atomic number of Ti = 22, Mn = 25, Fe = 26 and Co = 27)

- (A) $[Ti(H_2O)_6]^{2+}$
(B) $[Mn(CN)_6]^{3-}$
(C) $[Fe(CN)_6]^{3-}$
(D) $[Co(NH_3)_6]^{3+}$
-

83. A protein undergoes reversible thermal denaturation from its initial state N to denatured state D according to $N \rightleftharpoons D$. At $60^\circ C$, the concentrations of both N and D are equal at equilibrium, and the standard enthalpy change of denaturation is 666 kJ mol^{-1} . The standard entropy change (ΔS°) in $\text{kJ K}^{-1} \text{ mol}^{-1}$ of the protein upon denaturation at $60^\circ C$ is closest to

- (A) 11.1
(B) 2.0
(C) 2000.0
(D) 333.0
-

84. Given below are two statements: One is labelled as Assertion A and the other is labelled as Reason R.

Assertion A : Generally, 3d transition metals have high melting points.

Reason R : Involvement of 3d-electrons in addition to 4s-electrons in the interatomic metallic bonding.

In the light of the above statements, choose the most appropriate answer from the options given below.

- (A) A is not correct but R is correct.
(B) Both A and R are correct and R is the correct explanation of A.
(C) Both A and R are correct but R is NOT the correct explanation of A.
(D) A is correct but R is not correct.
-

85. Given below are two statements: One is labelled as Assertion A and the other is labelled as Reason R.

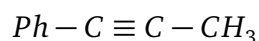
Assertion A: The first ionization enthalpy of O is lower than that of N and F

Reason R: The loss of an electron from O leads to stable half-filled p orbital.

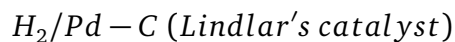
In light of the above statements, choose the most appropriate answer from the options given below:

- (A) A is not correct but R is correct
 - (B) Both A and R are correct and R is the correct explanation of A
 - (C) Both A and R are correct and R is NOT the correct explanation of A
 - (D) A is correct but R is not correct
-

86. Consider the following reaction sequences and choose the correct option.



On reduction with



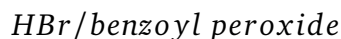
gives K.

On reduction with



gives L.

Further reaction with



gives M and N respectively.

- (A) M and N are stereoisomers
- (B) K and L are geometrical isomers
- (C) K and L are enantiomers
- (D) M and N are geometrical isomers

87. The highest occupied molecular orbital for Ne_2 is

- (A) σ_{2p}^*
 - (B) π_{2p}
 - (C) σ_{2p}
 - (D) π_{2p}^*
-

88. Match the species in List I with their geometry in List II

List I	List II
A. PCl_5 III	I. Tetrahedral
B. BrF_5	II. Square Planar
C. BF_4^- I	III. Trigonal bipyramidal
D. $[Ni(CN)_4]^{2-}$ II	IV. Square pyramidal

Choose the correct answer from the options given below:

- (1) A-III, B-II, C-I, D-IV
 - (2) A-IV, B-III, C-I, D-II
 - (3) A-III, B-IV, C-I, D-II
 - (4) A-III, B-I, C-II, D-IV
-

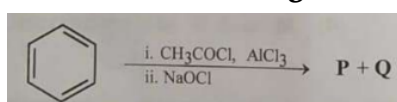
89. Match the vitamins in List I with their sources in List II

List I	List II
A. vitamin A _v	I. meat
B. vitamin B ₁₂	II. sunflower oil
C. vitamin E	III. green leafy vegetables
D. vitamin K _m	IV. carrots

Choose the correct answer from the options given below:

- (1) A-III, B-I, C-IV, D-II
 - (2) A-II, B-III, C-IV, D-I
 - (3) A-IV, B-I, C-II, D-III
 - (4) A-IV, B-II, C-I, D-III
-

90. For the following reaction sequence, choose the correct option



- (1) Both **P** and **Q** are carbonyl compounds.
 - (2) If **P** is the sodium salt of a carboxylic acid, **Q** is a primary alcohol.
 - (3) **P** and **Q** are aromatic compounds.
 - (4) If **P** gives a carboxylic acid on acidification, **Q** gives a poisonous gas on exposure to air and light.
-