

Question Paper

Conducted by the Central Board of Secondary Education (CBSE)



Time Allowed: 3 Hours

Maximum Marks: 70

General Instructions

- (i) This question paper contains **33 questions**. All questions are compulsory.
- (ii) Section A: Questions 1 to 16 are multiple choice questions, carrying 1 mark each.
- (iii) Section B: Questions 17 to 21 are very short answer questions, carrying 2 marks each.
- (iv) Section C: Questions 22 to 28 are short answer questions, carrying 3 marks each.
- (v) Section D: Questions 29 to 30 are case-based questions, carrying 4 marks each.
- (vi) Section E: Questions 31 to 33 are long answer questions, carrying 5 marks each.
- (vii) Attempt one alternative where an internal choice is provided. Follow the instructions within each question.
- (viii) Use of calculators is not allowed.

Section A

1. The molar ionic conductivities of Al^{3+} and SO_4^{2-} are $189 \text{ S cm}^2 \text{ mol}^{-1}$ and $160 \text{ S cm}^2 \text{ mol}^{-1}$ respectively. The value of limiting molar conductivity of $\text{Al}_2(\text{SO}_4)_3$ will be :

- (A) $198 \text{ S cm}^2 \text{ mol}^{-1}$ (B) $858 \text{ S cm}^2 \text{ mol}^{-1}$
(C) $588 \text{ S cm}^2 \text{ mol}^{-1}$ (D) $891 \text{ S cm}^2 \text{ mol}^{-1}$

2. Which of the following acids represents Vitamin C ?

- (A) Saccharic acid (B) Gluconic acid
(C) Ascorbic acid (D) Benzoic acid

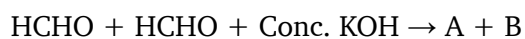
3. Rosenmund reduction is used for the preparation of Aldehydes. The catalyst used in this reaction is :

- (A) Pd-BaSO_4 (B) Anhydrous AlCl_3
(C) Iron (III) oxide (D) HgSO_4

4. From the elements of 3d series given below, which element shows the maximum number of oxidation states ?

- (A) Scandium (B) Manganese
(C) Chromium (D) Titanium

5. Consider the following reaction :



Identify A and B from the given options :

- (A) A – Methanol, B – Potassium formate (B) A – Ethanol, B – Potassium formate
(C) A – Methanal, B – Ethanol (D) A – Methanol, B – Potassium acetate

6. Which alkyl halide from the given options will undergo $\text{S}_{\text{N}}1$ reaction faster ?

- (A) $(\text{CH}_3)_3\text{C-Br}$ (B) $(\text{CH}_3)_2\text{CH-Br}$
(C) $\text{CH}_3\text{-CH}_2\text{-Br}$ (D) $(\text{CH}_3)_3\text{C-CH}_2\text{-Br}$

7. The fraction of molecules having energy equal to or greater than activation energy is :

- (A) A (B) $e^{-E_a/RT}$
(C) k (D) $A e^{-E_a/RT}$

8. Identify the primary amine from the given options :

- (A) $(\text{C}_2\text{H}_5)_3\text{N}$ (B) $(\text{C}_2\text{H}_5)_2\text{NH}$
(C) $\text{C}_2\text{H}_5\text{NH}_2$ (D) $(\text{CH}_3)_3\text{N}$

9. The general electronic configuration of d-block elements is :

- (A) $(n-1)d^{1-10}ns^{1-2}$ (B) $(n-1)d^1ns^{1-2}$
(C) $(n-1)d^1ns^{2-3}$ (D) $(n-1)d^0ns^{1-2}$

10. Match the reagents required for the given reactions :

I. Oxidation of primary alcohols to aldehydes (p) NaBH_4

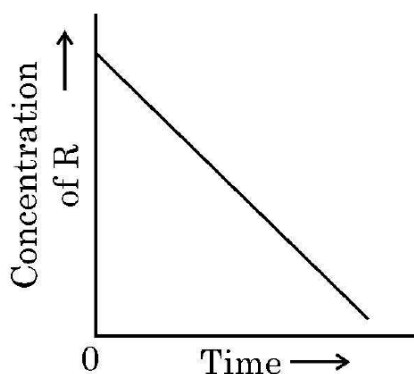
II. Butan-2-one to Butan-2-ol (q) 85% phosphoric acid at 440 K

III. Bromination of Phenol to 2, 4, 6-Tribromophenol (r) PCC

IV. Dehydration of propan-2-ol to propene (s) Bromine water

- (A) I-(r), II-(p), III-(s), IV-(q) (B) I-(q), II-(r), III-(p), IV-(s)
(C) I-(s), II-(q), III-(p), IV-(r) (D) I-(p), II-(s), III-(r), IV-(q)

11. In a given graph of zero order reaction (Concentration of R on Y-axis vs Time on X-axis), the slope and intercept are :



- (A) Slope = k , Intercept = $[R]_0$ (B) Slope = $-k$, Intercept = $[R]_0$
 (C) Slope = $k/2.303$, Intercept = $\ln[R]_0$ (D) Slope = $-k/2.303$, Intercept = $\ln A$

12. Nucleophilic addition of Grignard reagent to ketones followed by hydrolysis with dilute acids forms :

- (A) Alkene (B) Primary alcohol
 (C) Tertiary alcohol (D) Secondary alcohol

13. Two statements are given - one labelled Assertion (A) and the other labelled Reason (R).

Assertion (A) : Phenols are stronger acids than alcohols.

Reason (R) : Alkoxide ion is more stable than phenoxide ion.

- (A) Both Assertion (A) and Reason (R) are true and Reason (R) is the correct explanation of (A). (B) Both Assertion (A) and Reason (R) are true, but Reason (R) is not the correct explanation of (A).
 (C) Assertion (A) is true, but Reason (R) is false. (D) Assertion (A) is false, but Reason (R) is true.

14. **Assertion (A) :** For a Daniell cell, $Zn/Zn^{2+}(1M) || Cu^{2+}(1M)/Cu$ with $E^\circ_{cell} = 1.1 V$, if the external opposing potential is more than $1.1 V$, the electrons flow from Cu to Zn.

Reason (R) : Cell acts like a galvanic cell.

- (A) Both Assertion (A) and Reason (R) are true and Reason (R) is the correct explanation of (A). (B) Both Assertion (A) and Reason (R) are true, but Reason (R) is not the correct explanation of (A).
 (C) Assertion (A) is true, but Reason (R) is false. (D) Assertion (A) is false, but Reason (R) is true.

15. **Assertion (A) :** Benzoic acid does not undergo Friedel - Crafts reaction.

Reason (R) : Carboxyl group is deactivating and the catalyst aluminium chloride gets bonded to the carboxyl group.

- (A) Both Assertion (A) and Reason (R) are true and Reason (R) is the correct explanation of (A). (B) Both Assertion (A) and Reason (R) are true, but Reason (R) is not the correct explanation of (A).
 (C) Assertion (A) is true, but Reason (R) is false. (D) Assertion (A) is false, but Reason (R) is true.

16. **Assertion (A) :** Fructose is a reducing sugar.

Reason (R) : Fructose does not reduce Fehling solution and Tollen's reagent.

- (A) Both Assertion (A) and Reason (R) are true and Reason (R) is the correct explanation of (A). (B) Both Assertion (A) and Reason (R) are true, but Reason (R) is not the correct explanation of (A).
 (C) Assertion (A) is true, but Reason (R) is false. (D) Assertion (A) is false, but Reason (R) is true.

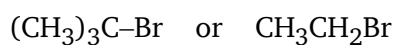
Section B

17. Define the following terms :

- (a) Half life period ($t_{1/2}$)
- (b) Effective collisions

18. A solution containing 60 g of a non-volatile solute in 250 g of water freezes at 270.67 K. Calculate the molar mass of the solute. (K_f of water = $1.86 \text{ K kg mol}^{-1}$)

19. (a) Which of the following compounds will react more rapidly by S_N1 reaction and why ?



(b) Arrange the following compounds in the increasing order of their boiling points :

Bromoform, Dibromomethane, Chloromethane, Bromomethane

20. (a) Write the stepwise mechanism of nucleophilic addition reactions in the carbonyl compounds.

OR

(b) How will you convert the following :

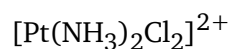
- (i) Toluene to benzoic acid.
- (ii) Ethanol to 3-Hydroxybutanal

21. (a) What happens when Glucose reacts with Nitric acid ? Write chemical equation.

(b) Write one structural difference between DNA and RNA.

Section C

22. (a) Draw the geometrical isomers of the given complex :



(b) Write the electronic configuration of d^5 ion if $\Delta_0 < P$

(c) What is an ambidentate ligand ?

23. Write chemical equations for the following reactions (Do any three) :

(a) Hydroboration - oxidation reaction

(b) Williamson Synthesis

(c) Friedel-Crafts Alkylation of Anisole

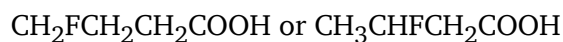
(d) Reimer-Tiemann Reaction

24. (a) Give chemical tests to distinguish between the following pairs of compounds :

(i) Phenol and Benzoic acid

(ii) Propanal and Propanone

(b) Which one of the given compounds is a stronger acid and why ?



25. Show that the time required for 99.9% completion in a first order reaction is 10 times of half-life ($t_{1/2}$) of the reaction. [$\log 2 = 0.3010$, $\log 10 = 1$]

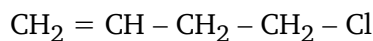
26. Define the following terms :

(a) Glycosidic linkage

(b) Primary structure of protein

(c) Disaccharides

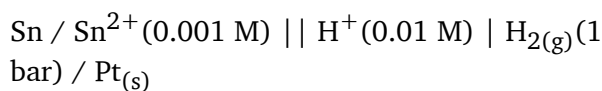
27. (a) Write the IUPAC name of the given compound :



(b) Why is thionyl chloride preferred for preparing alkyl halides from alcohols ?

(c) What happens when Methyl bromide reacts with KCN ?

28. Calculate emf of the following cell at 25 °C :



Given : $E^\circ(\text{Sn}^{2+}/\text{Sn}) = -0.14 \text{ V}$, $E^\circ(\text{H}^+/\text{H}_2) = 0.00 \text{ V}$ ($\log 10 = 1$)

29. Case Study : In a galvanic cell, chemical energy of a redox reaction is converted into electrical energy, whereas in an electrolytic cell the redox reaction occurs on passing electricity. The simplest galvanic cell is in which Zn rod is placed in a solution of ZnSO_4 and Cu rod is placed in a solution of CuSO_4 . The two rods are connected by a metallic wire through a voltmeter. The two solutions are joined by a salt bridge. The difference between the two electrode potentials of the two electrodes is known as electromotive force. In the process of electrolysis, the decomposition of a substance takes place by passing an electric current. One mole of electric charge when passed through a cell will discharge half a mole of a divalent metal ion such as Cu^{2+} . This was first formulated by Faraday in the form of laws of electrolysis.

Answer the following questions :

29.a. What is the function of a salt bridge in a galvanic cell ?

29.b. When does galvanic cell behave like an electrolytic cell ?

29.c. Can copper sulphate solution be stored in a pot made of zinc ? Explain with the help of the value of E°_{cell} .

$$(E^\circ \text{Cu}^{2+}/\text{Cu} = 0.34 \text{ V})$$

$$(E^\circ \text{Zn}^{2+}/\text{Zn} = -0.76 \text{ V})$$

OR

How much charge in terms of Faraday is required for the following :

(i) 1 mol of MnO_4^- to Mn^{2+}

(ii) 1 mol of H_2O to O_2

30. Case Study : The nature of bonding, structure of the coordination compound can be explained to some extent by valence bond theory. The central metal atom/ion makes available a number of vacant orbitals equal to its coordination number. The appropriate atomic orbitals (s, p and d) of the metal hybridise to give a set of equivalent orbitals of definite geometry such as square planar, tetrahedral, octahedral and so on. A strong covalent bond is formed only when the orbitals overlap to the maximum extent. The d-orbitals involved in the hybridisation may be either inner d-orbitals i.e. $(n-1)d$ or outer d-orbitals i.e. nd . The complexes formed are called inner orbital complex (low spin complex) and outer orbital complex (high spin complex) respectively. Further, the complexes can be paramagnetic or diamagnetic in nature. The drawbacks of this theory are that this involves number of assumptions and also does not explain the colour of the complex.

Answer the following questions :

30.a. Predict whether $[\text{CoF}_6]^{3-}$ is diamagnetic or paramagnetic and why ? [Atomic number : Co = 27]

- | | |
|--|---|
| (A) Paramagnetic; F^- is a weak field ligand so Co^{3+} ($3d^6$) remains high-spin with 4 unpaired electrons (sp^3d^2 , outer orbital). | (B) Diamagnetic; F^- is a strong field ligand and forces pairing in Co^{3+} . |
| (C) Paramagnetic; with 2 unpaired electrons because Co^{3+} uses d^2sp^3 hybridisation. | (D) Diamagnetic; because Co^{3+} has fully paired 3d electrons in any ligand field. |

30.b. What is the coordination number of Co in $[\text{Co}(\text{en})_2\text{Cl}_2]^+$?

- | | |
|-------|-------|
| (A) 4 | (B) 5 |
| (C) 6 | (D) 8 |

30.c. (i) Write the IUPAC name of the given complex : $[\text{Pt}(\text{NH}_3)_2\text{Cl}_2]^{2+}$

(ii) Explain whether $[\text{Co}(\text{NH}_3)_6]^{3+}$ is an inner orbital or outer orbital complex.

OR

Using valence bond theory, deduce the shape and hybridisation of $[\text{Ni}(\text{NH}_3)_6]^{2+}$. [Atomic number of Ni = 28]

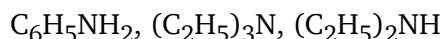
31. (a) An amide 'A' with molecular formula C_7H_7ON undergoes Hoffmann Bromamide degradation reaction to give amine 'B'. 'B' on treatment with nitrous acid at 273-278 K forms 'C' and on treatment with chloroform and ethanolic potassium hydroxide forms 'D'. 'C' on treatment with ethanol gives 'E'. Identify 'A', 'B', 'C', 'D' and 'E' and write the sequence of chemical equations.

OR

(b)

(i) (1) What is Hinsberg's reagent ?

(2) Arrange the following compounds in the increasing order of their basic strength in gaseous phase :



(ii) Give reasons for the following :

(1) Methyl amine is more basic than aniline.

(2) Aniline readily reacts with bromine water to give 2, 4, 6-tribromoaniline.

(3) Primary amines have higher boiling points than tertiary amines.

32. Attempt any five of the following :

(a) Why Zinc is not regarded as a transition element ?

(b) What is Lanthanoid contraction ?

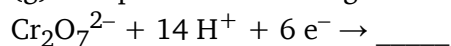
(c) Why is first ionization enthalpy of chromium lower than that of Zn ?

(d) Why are transition elements good catalysts ?

(e) Compounds of transition metals are generally coloured. Give reason.

(f) Out of $KMnO_4$ and K_2MnO_4 , which one is paramagnetic and why ?

(g) Complete the following ionic equation :



33. (a)

(i) Define reverse osmosis.

(ii) Why are aquatic species more comfortable in cold water in comparison to warm water ?

(iii) A solution containing 2 g of glucose ($M = 180 \text{ g mol}^{-1}$) in 100 g of water is prepared at 303 K. If the vapour pressure of pure water at 303 K is 32.8 mm Hg, what would be the vapour pressure of the solution ?

OR

(b)

(i) Predict whether Van't Hoff factor will be less or greater than one, when Ethanoic acid is dissolved in benzene.

(ii) Define ideal solution.

(iii) Calculate the mass of $CaCl_2$ (molar mass = 111 g mol^{-1}) to be dissolved in 500 g of water to lower its freezing point by 2 K, assuming that $CaCl_2$ undergoes complete dissociation. (K_f for water = $1.86 \text{ K kg mol}^{-1}$)

- End of Question Paper -