CBSE Class 12 Chemistry 2026 - 56-5-1

Question Paper

Conducted by the Central Board of Secondary Education (CBSE)

Time Allowed: 3 Hours **Maximum Marks: 70**

General Instructions

(i) This question paper contains **33 questions**. All questions are compulsory.

(ii) Section A: Questions 1 to 16 are multiple choice questions, carrying 1 mark each.

(iii) Section B: Questions 17 to 21 are very short answer questions, carrying 2 marks each.

(iv) Section C: Questions 22 to 28 are short answer questions, carrying 3 marks each.

(v) Section D: Questions 29 to 30 are case-based questions, carrying 4 marks each.

(vi) Section E: Questions 31 to 33 are long answer questions, carrying 5 marks each.

(vii) Attempt one alternative where an internal choice is provided. Follow the instructions within each question.

(viii) Use of calculators is not allowed.

Section A

1. Which of the reactions is used in the conversion of a ketone into hydrocarbon ?

(A) Reimer-Tiemann reaction
(B) Wolff-Kishner reduction
(C) Aldol condensation
(D) Stephen reaction

2. Which of the following reagents are used to prepare primary amines by Hoffmann bromamide degradation reaction ?

(i) $R\text{-}CO\text{-}NH_2$ (ii) NaOH (iii) Br_2 (iv) $CHCl_3$

(A) (i), (ii) and (iv)
(B) (i) and (iii)
(C) (i), (ii) and (iii)
(D) (i), (iii) and (iv)

3. The major product of carbylamine reaction is :

(A) Carboxylic acid
(B) Aldehyde
(C) Cyanide
(D) Isocyanide

4. Actinoids show larger number of oxidation states :

(A) because they are radioactive in nature
(B) because they have large atomic numbers
(C) because they have large atomic masses
(D) due to comparable energies of 5f, 6d and 7s orbitals

5. Consider the following reaction and identify A and B :

$CH_3Cl + NaI \rightarrow$ (dry acetone) $A + B$

(A) $A = CH_3I$, $B = NaCl$
(B) $A = CH_3OH$, $B = NaCl$
(C) $A = CH_3CHO$, $B = NaCl$
(D) $A = C_2H_6$, $B = CH_3I$

6. The correct formula of Hinsberg's reagent is :

(A) C_6H_5COCl
(B) $C_6H_5SO_2Cl$
(C) $C_6H_5CONHCH_3$
(D) $C_6H_5CH_2NH_2$

7. Half-life ($t_{1/2}$) of a first order reaction is 1386 s. The value of rate constant is :

(A) $0.5 \times 10^{-4}\ s^{-1}$
(B) $5.0 \times 10^{-4}\ s^{-1}$
(C) $0.5 \times 10^{-5}\ s^{-1}$
(D) $0.5 \times 10^{-2}\ s^{-1}$

8. Which of the following ligands forms a chelate complex ?

(A) Ammonia
(B) Water
(C) NO_2^-
(D) Oxalate ion

9. Primary, secondary and tertiary alcohols can be distinguished by :

(A) Lucas test
(B) Fehling's test
(C) Tollens' test
(D) Hinsberg's test

10. Consider the following compounds :

$C_2H_5NH_2$, $(C_2H_5)_2NH$, $C_6H_5CH_2NH_2$, NH_3, $C_6H_5NH_2$

The correct increasing order of the above compounds on the basis of their basic strength is :

(A) $C_6H_5NH_2 < NH_3 < C_6H_5CH_2NH_2 < C_2H_5NH_2 < (C_2H_5)_2NH$
(B) $NH_3 < C_6H_5NH_2 < C_6H_5CH_2NH_2 < C_2H_5NH_2 < (C_2H_5)_2NH$
(C) $C_6H_5CH_2NH_2 < (C_2H_5)_2NH < NH_3 < C_6H_5NH_2 < C_2H_5NH_2$
(D) $C_2H_5NH_2 < C_6H_5NH_2 < NH_3 < C_6H_5CH_2NH_2 < (C_2H_5)_2NH$

11. Identify the polysaccharide among the following :

(A) Fructose
(B) Maltose
(C) Glucose
(D) Cellulose

12. The polypeptide chain in a protein has amino acids linked with each other in a specific sequence. This specific sequence of amino acids is called :

(A) Primary structure of protein
(B) Secondary structure of protein
(C) Tertiary structure of protein
(D) Quaternary structure of protein

13. **Assertion (A):** D(+)-Glucose is dextrorotatory in nature.

Reason (R): (+) represents dextrorotatory nature and D represents the configuration.

(A) Both Assertion (A) and Reason (R) are true and Reason (R) is the correct explanation of the Assertion (A).
(B) Both Assertion (A) and Reason (R) are true, but Reason (R) is not the correct explanation of the Assertion (A).
(C) Assertion (A) is true, but Reason (R) is false.
(D) Assertion (A) is false, but Reason (R) is true.

14. **Assertion (A):** Highest oxidation state of Mn is +7 in first series of transition elements.

Reason (R): Transition metals exhibit variable oxidation states.

(A) Both Assertion (A) and Reason (R) are true and Reason (R) is the correct explanation of the Assertion (A).
(B) Both Assertion (A) and Reason (R) are true, but Reason (R) is not the correct explanation of the Assertion (A).
(C) Assertion (A) is true, but Reason (R) is false.
(D) Assertion (A) is false, but Reason (R) is true.

15. **Assertion (A):** p-nitrophenol is more acidic than phenol.

Reason (R): Nitro group is an electron-withdrawing group, it stabilises phenoxide ion by dispersal of negative charge.

(A) Both Assertion (A) and Reason (R) are true and Reason (R) is the correct explanation of the Assertion (A).
(B) Both Assertion (A) and Reason (R) are true, but Reason (R) is not the correct explanation of the Assertion (A).
(C) Assertion (A) is true, but Reason (R) is false.
(D) Assertion (A) is false, but Reason (R) is true.

16. **Assertion (A):** All aliphatic aldehydes give a positive Fehling's test.

Reason (R): Aliphatic aldehydes are reduced by Fehling's reagent.

(A) Both Assertion (A) and Reason (R) are true and Reason (R) is the correct explanation of the Assertion (A).
(B) Both Assertion (A) and Reason (R) are true, but Reason (R) is not the correct explanation of the Assertion (A).
(C) Assertion (A) is true, but Reason (R) is false.
(D) Assertion (A) is false, but Reason (R) is true.

Section B

17. (a) 1.00 molal aqueous solution of trichloroacetic acid is heated to its boiling point. Boiling point of this solution was found to be 100.18°C. Calculate the van't Hoff factor for trichloroacetic acid.

(Given : K_b for water = 0.512 K kg mol^{-1})

OR

(b) State Henry's law. Calculate the mole fraction of CO_2 in water at 298 K under 760 mm Hg.

(Given : K_H for CO_2 in H_2O at 298 K = 1.25×10^6 mm Hg)

18. (a) Name the cell which was used in the Apollo space programme for providing electrical power.

(b) Define limiting molar conductivity.

19. (a) Complete the following equation :

$C_6H_5\text{-}N_2^+Cl^- + C_6H_5\text{-}OH \rightarrow$

(b) How will you convert nitromethane to methyl isocyanide ?

20. (a) What are the products obtained on hydrolysis of sucrose ?

(b) What are essential amino acids ?

21. (a) Write any two fat soluble vitamins.

(b) How will you confirm the presence of five -OH groups in a glucose molecule, which are attached to different carbon atoms ?

Section C

22. Calculate emf of the following cell at 298 K :

$Cr(s) \mid Cr^{3+}(aq)\ (0.1\ M) \mid\mid Fe^{2+}(aq)\ (0.01\ M) \mid Fe(s)$

(Given : $E^{\circ}_{Cr^{3+}/Cr} = -0.74$ V, $E^{\circ}_{Fe^{2+}/Fe} = -0.44$ V, [log 10 = 1])

23. (a) Define order of a reaction.

(b) The rate for the following reaction is given by : $A + B \rightarrow C$, Rate $= k[A][B]^2$

(i) How is the rate of reaction affected if we double the concentration of B ?

(ii) Write the overall order of a reaction if 'A' is present in large excess.

24. The rate of the chemical reaction doubles when the temperature is raised from 298 K to 308 K. Calculate activation energy (E_a) for this reaction assuming that it does not change with temperature.

(Given : R = 8.314 J mol^{-1} K^{-1}, log 2 = 0.30)

25. (a) Write the IUPAC name of the following complex : $K_3[Cr(C_2O_4)_3]$

(b) Differentiate between homoleptic complex and heteroleptic complex.

(c) Which type of isomerism is exhibited by the following complex : $[Pt(NH_3)_2Cl_2]$

26. (a) A coordination compound $CrCl_3.6H_2O$ is mixed with excess of $AgNO_3$ solution, two moles of AgCl are precipitated per mole of the compound. Write the structural formula of the coordination compound.

(b) Write the oxidation state and hybridisation of the central metal in the following complex : $[Fe(H_2O)_6]^{3+}$ [Atomic number of Fe = 26]

(c) Why is $[Ni(H_2O)_6]^{2+}$ coloured ? [Atomic number of Ni = 28]

27. How do you convert the following ?

(a) Acetophenone to Benzoic acid

(b) Acetonitrile to Acetone

(c) Benzoic acid to Benzene

28. (a) (i) Arrange the following compounds in increasing order of their acidic strengths : $CH_3CH(Br)CH_2COOH$, CH_3-$CH(CH_3)$-COOH, $CH_3CH_2CH(Br)COOH$

(ii) Why is CH_3CHO more reactive than acetone towards reaction with HCN ?

(iii) Complete the equation : $(CH_3)(H)C{=}O + NH_2$-$NH_2 \rightarrow (H^+)$

OR

(b) An organic compound with the molecular formula C_8H_8O forms 2,4-DNP derivative, reduces Tollens' reagent and undergoes Cannizzaro reaction. On vigorous oxidation it gives Benzene-1,2-dicarboxylic acid. Identify the compound and write the reactions of the compound with 2,4-DNP and when it undergoes Cannizzaro reaction.

Section D

29. **Case Study:** Ethers are prepared by the dehydration of alcohols in presence of protic acids at 413 K. Symmetrical and unsymmetrical ethers can also be prepared by Williamson synthesis. This reaction involves S_N2 attack of an alkoxide ion on primary alkyl halide. If a tertiary alkyl halide is used, elimination reaction occurs and an alkene is formed, no ether is formed. C-O bond in ethers are cleaved under drastic conditions with excess of HI. When unsymmetrical ethers react with HI, the alkyl halide is formed from smaller alkyl group. If one of the alkyl group is tertiary, the alkyl halide is formed from the tertiary alkyl group because tertiary alkyl carbocation is more stable than the primary carbocation. Cleavage of alkyl aryl ethers takes place at the alkyl-oxygen bond due to more stable aryl-oxygen bond. The order of reactivity of hydrogen halides is HI > HBr > HCl. Aromatic ethers undergo electrophilic substitution reactions. The alkoxy group attached to the aromatic ring activates the ring towards electrophilic substitution and directs the incoming group to ortho- and para-positions.

Answer the following questions :

(a) Complete the following equations :

(i) $C_6H_5\text{-}OCH_3 + CH_3Cl \rightarrow$ (anhyd. $AlCl_3$)

(ii) $C_6H_5\text{-}OCH_3$ + conc. ($HNO_3 + H_2SO_4$) $\rightarrow$

(b)(i) Write the names of alkyl halide and sodium alkoxide used to prepare tert-butyl ethyl ether.

OR

(b)(ii) Anisole on reaction with HI gives phenol and CH_3-I and not methanol and iodobenzene. Justify the statement.

(c) Why is C-O-C bond angle in ethers slightly greater than tetrahedral angle ?

30. **Case Study:** Electrochemistry is the study of the relationship between chemical energy and electrical energy. Many spontaneously occurring chemical reactions liberate electrical energy. In electrolysis, electrical energy is converted directly into chemical energy. The product of an electrolytic reaction depends on the nature of the material being electrolysed and the type of electrode used. Oxidising and reducing species present in the electrolytic cell and their standard electrode potential too, affect the products of electrolysis. Electrolysis plays an important role in most people's daily lives, whether it is for the manufacturing of aluminium, electroplating of metals, or the synthesis of chemical compounds. Michael Faraday was the first scientist who proposed two laws to explain the quantitative aspects of electrolysis, popularly known as Faraday's laws of electrolysis. Faraday's laws of electrolysis provide a basis for mathematical analysis of the mass deposited at electrodes and the amount of charge passed through them. Faraday's laws are fundamental in various applications, including electroplating, metal extraction, battery technology and chemical synthesis. These laws also help in environmental monitoring and in various chemistry experiments.

30. (continued)

Answer the following questions :

(a) Predict the products of electrolysis in each of the following :

(i) An aqueous solution of $CuCl_2$ with platinum electrodes.

(ii) A concentrated solution of H_2SO_4 with platinum electrodes.

(b)(i) How much charge in faraday is required for the reduction of 1 mol of Ag^+ to Ag ?

OR

(b)(i) State Faraday's second law of electrolysis.

(c) The following reactions occur at the anode during the electrolysis of aqueous sodium chloride solution :

(I) $Cl^-(aq) \rightarrow \frac{1}{2} Cl_2(g) + e^-$, $E^\circ_{(cell)} = 1.36$ V

(II) $2H_2O(l) \rightarrow O_2(g) + 4H^+(aq) + 4e^-$, $E^\circ_{(cell)} = 1.23$ V

Which reaction is feasible at the anode and why ?

Section E

31. (a)(i) Calculate the freezing point of a solution when 10.5 g of $MgBr_2$ was dissolved in 250 g of water, assuming $MgBr_2$ undergoes complete dissociation.

(Given : Molar mass of $MgBr_2$ = 184 g mol^{-1}, K_f for water = 1.86 K kg mol^{-1})

(ii) Write two differences between ideal and non-ideal solutions.

OR

(b)(i) A solution is prepared by dissolving 0.025 g of potassium sulphate in 2 L of water at 27°C. Assuming potassium sulphate is completely dissociated, determine its osmotic pressure.

(Given : R = 0.082 L atm K^{-1} mol^{-1}, Molar mass of K_2SO_4 = 174 g mol^{-1})

(ii) What type of azeotrope will be formed by a solution of acetone and chloroform ? Give reason.

32. (a)(i) (I) Why do transition metals show variable oxidation states ?

(II) Out of Mn^{2+} and Ti^{2+} which will be more paramagnetic and why ? [Atomic No. : Ti = 22, Mn = 25]

(III) Which ion is the strongest oxidising agent in the options given below : Cr^{3+}, V^{3+}, Mn^{3+}. Give reason. [Atomic No. : Cr = 24, V = 23, Mn = 25]

(ii) Complete and balance the following equations :

(I) $2MnO_2 + 4KOH + O_2 \rightarrow$

(II) $5C_2O_4^{2-} + 2MnO_4^- + 16H^+ \rightarrow$

OR

(b)(i) What is meant by lanthanoid contraction ?

(ii) Why do transition metals form coloured compounds ?

(iii) Why are $E^\circ_{M^{2+}/M}$ values for Mn and Zn more negative than expected ?

(iv) Which is the most stable oxidation state of Cu and why ?

(v) Why is Ce^{4+} in aqueous solution a good oxidising agent ?

33. (a)(i) (I) Which of the following is more reactive towards S_N1 reaction : 2-Bromo-2-methylbutane or 1-Bromopentane

(II) What type of halide is present in the following compound : CH_3-$CH(CH_3)$-$C(Cl)=CH_2$

(III) Why is chloroform stored in dark coloured bottles ?

(ii) Define the following terms :

(I) Ambident Nucleophiles (II) Racemic mixture

OR

(b)(i) Answer the following :

(I) Which isomer of C_4H_9Br is most reactive towards S_N1 reaction ?

(II) Predict the alkene that would be formed by dehydrohalogenation of 1-Bromo-1-methylcyclohexane.

(III) Although chlorine shows strong -I effect, yet it is ortho/para-directing in electrophilic aromatic substitution reactions. Why ?

(ii) Write the major product in the following reactions :

(I) 2 [4-isopropylchlorobenzene] + 2 Na $\rightarrow$ (dry ether)

(II) 4-nitrochlorobenzene + Br_2 $\rightarrow$ (heat)

- End of Question Paper -